

NAME (Print): _____

Chemistry 320N
2nd Midterm Exam
March 13, 2025

EID _____

SIGNATURE: _____

**Please print the
first three letters
of your last name
in the three boxes**

--	--	--

Please Note: Please take your time. You have three hours to take this exam. Please do not rush, we want you to show us everything you have learned this semester so far! Making careless mistakes is not good for anyone! If you find yourself getting anxious because of a problem, skip it and come back. Please do not second guess yourself! Keep track of the questions worth a lot of points. (This does not mean they are hard, it just means we think they cover important material.)

One last thing: I recommend you close your eyes for a moment, then take some nice deep breaths before you begin. **YOU GOT THIS!**

FINALLY, DUE TO SOME UNFORTUNATE RECENT INCIDENTS YOU ARE NOT ALLOWED TO INTERACT WITH YOUR CELL PHONE IN ANY WAY. IF YOU TOUCH YOUR CELL PHONE DURING THE EXAM YOU WILL GET A "0" NO MATTER WHAT YOU ARE DOING WITH THE PHONE. PUT IT AWAY AND LEAVE IT THERE!!!

Student Honor Code

"As a student of The University of Texas at Austin, I shall abide by the core values of the University and uphold academic integrity."

(Your signature)

PERIODIC TABLE OF THE ELEMENTS

Elementary Subatomic Particles

	Electron	Proton	Neutron	Photon	Neutrino
Symbol	e	p	n	γ	ν
Rest mass (kg)	9.109382(1) × 10 ⁻³¹	1.672621(1) × 10 ⁻²⁷	1.674928(1) × 10 ⁻²⁷	0	~0
Major mass (kg/mol)	5.485799(1) × 10 ⁻⁷	1.0072764(1) × 10 ⁻³	1.0086649(1) × 10 ⁻³	0	~0
Particle-electron mass ratio	1	1836.1527(1) × 10 ⁻³	1838.68036(1) × 10 ⁻³	0	~0
Particle-proton mass ratio	5.4461701(1) × 10 ⁻⁴	1	1.00137846(1) × 10 ⁻³	0	~0
Particle-neutron mass ratio	5.438672(1) × 10 ⁻⁴	0.99904	1	0	~0
Spin quantum number (h/2π)	-1/2, 1/2, 3/2, 5/2, ...	1/2, 3/2, 5/2, ...	1/2, 3/2, 5/2, ...	0	0
Spin (h)	~1 × 10 ⁻³⁵	~1 × 10 ⁻³⁵	~1 × 10 ⁻³⁵	0	0
Spin quantum number	1/2	1/2	1/2	0	1/2
Compton wavelength (m)	2.42631024(1) × 10 ⁻¹²	1.32141002(1) × 10 ⁻¹³	1.31959110(1) × 10 ⁻¹³	~	~
Magnetic moment (μ _B)	9.2847637(1) × 10 ⁻²⁴	1.81887903(1) × 10 ⁻²³	0.9818247(1) × 10 ⁻²³	0	0
In their magnitudes: μ _B	1.0011586519(1) × 10 ⁻²³	1.81887903(1) × 10 ⁻²³	0.9818247(1) × 10 ⁻²³	0	0
In their magnitudes: μ _N	1.76192726(1) × 10 ⁻²⁷	1.76192726(1) × 10 ⁻²⁷	1.76192726(1) × 10 ⁻²⁷	0	0

Summary particles are the fundamental constituents of energy and matter. The electron (e⁻) is a negative-energy particle which has the same mass as an ordinary electron. The neutrino (ν) has neither positive nor negative charge. It is thought to spin in the direction of motion. Neutrino-like charge is the transformation of a neutrino into a proton, a beta particle (negative electron) and a positron (e⁺) in a positive-energy particle which has the same mass as an ordinary electron. The antineutrino (ν̄) has neither positive nor negative charge. It is thought to spin in the direction of motion. Neutrino-like charge is the transformation of a neutrino into a proton, a beta particle (negative electron) and a positron (e⁺) in a positive-energy particle which has the same mass as an ordinary electron. The neutrino (ν) has neither positive nor negative charge. It is thought to spin in the direction of motion. Neutrino-like charge is the transformation of a neutrino into a proton, a beta particle (negative electron) and a positron (e⁺) in a positive-energy particle which has the same mass as an ordinary electron. The antineutrino (ν̄) has neither positive nor negative charge. It is thought to spin in the direction of motion. Neutrino-like charge is the transformation of a neutrino into a proton, a beta particle (negative electron) and a positron (e⁺) in a positive-energy particle which has the same mass as an ordinary electron.

% Ionic Character of a Single Chemical Bond

Percent ionic character describes the nature of a bond. Bonds possessing 50% or greater ionic character are commonly termed ionic. Bonds with less than 50% ionic character are termed covalent. Pauling's equation was modified by Harnay & Smith in order to achieve better agreement between experimental and calculated values. Transition from covalent to ionic bonding is usually accompanied by a reduction in electrical conductivity, melting point and boiling point.

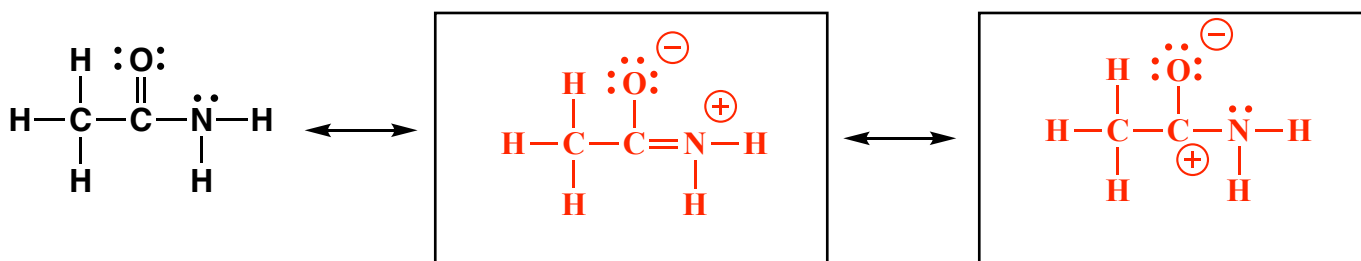
1A		2A										3A										4A										5A										6A										7A										8A																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																													
IA		IIA										IIIA										IVA										VA										VIA										VIIA										VIIIA																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																													
1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36	37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54	55	56	57	58	59	60	61	62	63	64	65	66	67	68	69	70	71	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86	87	88	89	90	91	92	93	94	95	96	97	98	99	100	101	102	103	104	105	106	107	108	109	110	111	112	113	114	115	116	117	118	119	120	121	122	123	124	125	126	127	128	129	130	131	132	133	134	135	136	137	138	139	140	141	142	143	144	145	146	147	148	149	150	151	152	153	154	155	156	157	158	159	160	161	162	163	164	165	166	167	168	169	170	171	172	173	174	175	176	177	178	179	180	181	182	183	184	185	186	187	188	189	190	191	192	193	194	195	196	197	198	199	200	201	202	203	204	205	206	207	208	209	210	211	212	213	214	215	216	217	218	219	220	221	222	223	224	225	226	227	228	229	230	231	232	233	234	235	236	237	238	239	240	241	242	243	244	245	246	247	248	249	250	251	252	253	254	255	256	257	258	259	260	261	262	263	264	265	266	267	268	269	270	271	272	273	274	275	276	277	278	279	280	281	282	283	284	285	286	287	288	289	290	291	292	293	294	295	296	297	298	299	300	301	302	303	304	305	306	307	308	309	310	311	312	313	314	315	316	317	318	319	320	321	322	323	324	325	326	327	328	329	330	331	332	333	334	335	336	337	338	339	340	341	342	343	344	345	346	347	348	349	350	351	352	353	354	355	356	357	358	359	360	361	362	363	364	365	366	367	368	369	370	371	372	373	374	375	376	377	378	379	380	381	382	383	384	385	386	387	388	389	390	391	392	393	394	395	396	397	398	399	400	401	402	403	404	405	406	407	408	409	410	411	412	413	414	415	416	417	418	419	420	421	422	423	424	425	426	427	428	429	430	431	432	433	434	435	436	437	438	439	440	441	442	443	444	445	446	447	448	449	450	451	452	453	454	455	456	457	458	459	460	461	462	463	464	465	466	467	468	469	470	471	472	473	474	475	476	477	478	479	480	481	482	483	484	485	486	487	488	489	490	491	492	493	494	495	496	497	498	499	500	501	502	503	504	505	506	507	508	509	510	511	512	513	514	515	516	517	518	519	520	521	522	523	524	525	526	527	528	529	530	531	532	533	534	535	536	537	538	539	540	541	542	543	544	545	546	547	548	549	550	551	552	553	554	555	556	557	558	559	560	561	562	563	564	565	566	567	568	569	570	571	572	573	574	575	576	577	578	579	580	581	582	583	584	585	586	587	588	589	590	591	592	593	594	595	596	597	598	599	600	601	602	603	604	605	606	607	608	609	610	611	612	613	614	615	616	617	618	619	620	621	622	623	624	625	626	627	628	629	630	631	632	633	634	635	636	637	638	639	640	641	642	643	644	645	646	647	648	649	650	651	652	653	654	655	656	657	658	659	660	661	662	663	664	665	666	667	668	669	670	671	672	673	674	675	676	677	678	679	680	681	682	683	684	685	686	687	688	689	690	691	692	693	694	695	696	697	698	699	700	701	702	703	704	705	706	707	708	709	710	711	712	713	714	715	716	717	718	719	720	721	722	723	724	725	726	727	728	729	730	731	732	733	734	735	736	737	738	739	740	741	742	743	744	745	746	747	748	749	750	751	752	753	754	755	756	757	758	759	760	761	762	763	764	765	766	767	768	769	770	771	772	773	774	775	776	777	778	779	780	781	782	783	784	785	786	787	788	789	790	791	792	793	794	795	796	797	798	799	800	801	802	803	804	805	806	807	808	809	810	811	812	813	814	815	816	817	818	819	820	821	822	823	824	825	826	827	828	829	830	831	832	833	834	835	836	837	838	839	840	841	842	843	844	845	846	847	848	849	850	851	852	853	854	855	856	857	858	859	860	861	862	863	864	865	866	867	868	869	870	871	872	873	874	875	876	877	878	879	880	881	882	883	884	885	886	887	888	889	890	891	892	893	894	895	896	897	898	899	900	901	902	903	904	905	906	907	908	909	910	911	912	913	914	915	916	917	918	919	920	921	922	923	924	925	926	927	928	929	930	931	932	933	934	935	936	937	938	939	940	941	942	943	944	945	946	947	948	949	950	951	952	953	954	955	956	957	958	959	960	961	962	963	964	965	966	967	968	969	970	971	972	973	974	975	976	977	978	979	980	981	982	983	984	985	986	987	988	989	990	991	992	993	994	995	996	997	998	999	1000	1001	1002	1003	1004	1005	1006	1007	1008	1009	1010	1011	1012	1013	1014	1015	1016	1017	1018	1019	1020	1021	1022	1023	1024	1025	1026	1027	1028	1029	1030	1031	1032	1033	1034	1035	1036	1037	1038	1039	1040	1041	1042	1043	1044	1045	1046	1047	1048	1049	1050	1051	1052	1053	1054	1055	1056	1057	1058	1059	1060	1061	1062	1063	1064	1065	1066	1067	1068	1069	1070	1071	1072	1073	1074	1075	1076	1077	1078	1079	1080	1081	1082	1083	1084	1085	1086	1087	1088	1089	1090	1091	1092	1093	1094	1095	1096	1097	1098	1099	1100	1101	1102	1103	1104	1105	1106	1107	1108	1109	1110	1111	1112	1113	1114	1115	1116	1117	1118	1119	1120	1121	1122	1123	1124	1125	1126	1127	1128	1129	1130	1131	1132	1133	1134	1135	1136	1137	1138	1139	1140	1141	1142	1143	1144	1145	1146	1147	1148	1149	1150	1151	1152	1153	1154	1155	1156	1157	1158	1159	1160	1161	1162	1163	1164	1165	1166	1167	1168	1169	1170	1171	1172	1173	1174	1175	1176	1177	1178	1179	1180	1181	1182	1183	1184	1185	1186	1187	1188	1189	1190	1191	1192	1193	1194	1195	1196	1197	1198	1199	1200	1201	1202	1203	1204	1205	1206	1207	1208	1209	1210	1211	1212	1213	1214	1215	1216	1217	1218	1219	1220	1221	1222	1223	1224	1225	1226	1227	

Compound		pK _a
Hydrochloric acid	H-Cl	-7
Protonated alcohol	$\text{RCH}_2\text{OH}_2^+$	-2
Hydronium ion	H_3O^+	-1.7
Carboxylic acids	$\text{R}-\overset{\text{O}}{\parallel}{\text{C}}-\text{H}$	3-5
Thiols	RCH_2SH	8-9
Ammonium ion	H_4N^+	9.2
β-Dicarbonyls	$\text{RC}-\overset{\text{O}}{\parallel}{\text{C}}-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{R}'$	10
Primary ammonium	$\text{H}_3\text{N}^+\text{CH}_2\text{CH}_3$	10.5
β-Ketoesters	$\text{RC}-\overset{\text{O}}{\parallel}{\text{C}}-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{OR}'$	11
β-Diesters	$\text{ROC}-\overset{\text{O}}{\parallel}{\text{C}}-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{OR}'$	13
Water	HOH	15.7
Alcohols	RCH_2OH	15-19
Acid chlorides	$\text{RCH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{Cl}$	16
Aldehydes	$\text{RCH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{H}$	18-20
Ketones	$\text{RCH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{R}'$	18-20
Esters	$\text{RCH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{OR}'$	23-25
Terminal alkynes	$\text{RC}\equiv\text{C}-\text{H}$	25
LDA	$\text{H}-\text{N}(\text{i-C}_3\text{H}_7)_2$	40
Terminal alkenes	$\text{R}_2\text{C}=\underset{\text{H}}{\text{C}}-\text{H}$	44
Alkanes	$\text{CH}_3\text{CH}_2-\text{H}$	51

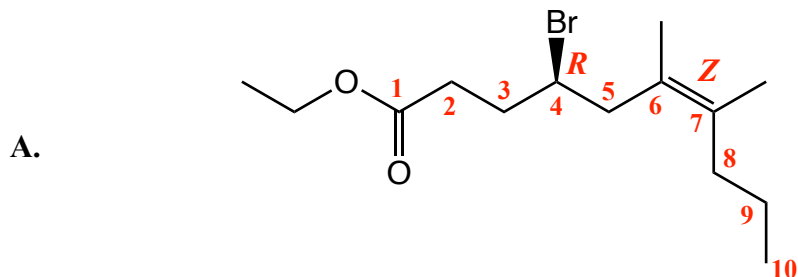
1. (5 pts) What is the most important question in organic chemistry?

Where are the electrons?

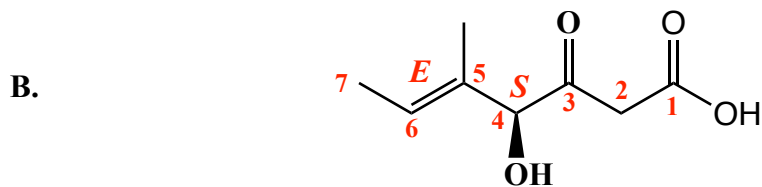
2. (10 pts) Amides are best represented as the hybrid of three contributing structures. Draw the second and third important contributing structures in the spaces provided. (No need to draw any arrows for this.)



3. (6 pts each) Write an acceptable IUPAC name or draw a structural formula for the following molecules:



**ethyl (*R,Z*)-4-bromo-6,7-dimethyldec-6-enoate
or ethyl (*R,Z*)-4-bromo-6,7-dimethyl-6-decenoate**

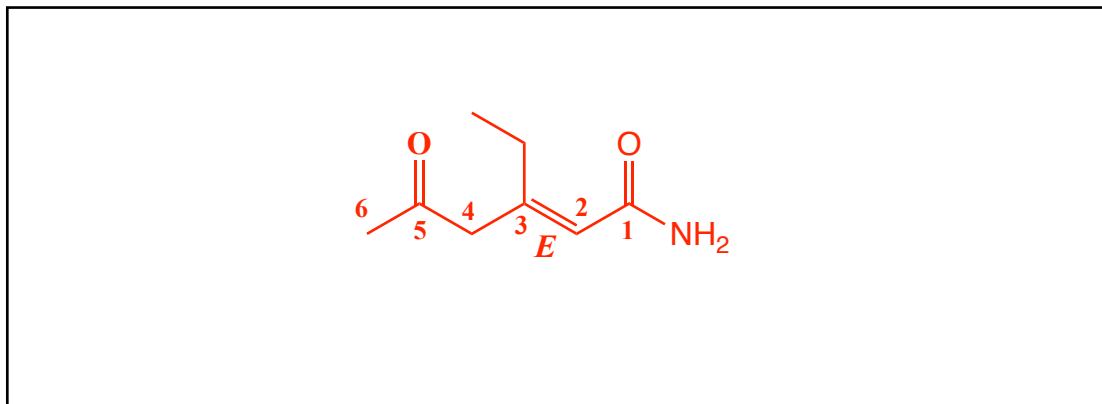


**(*S,E*)-4-hydroxy-5-methyl-3-oxohept-5-enoic acid
or (*S,E*)-4-hydroxy-5-methyl-3-oxo-5-heptenoic acid**

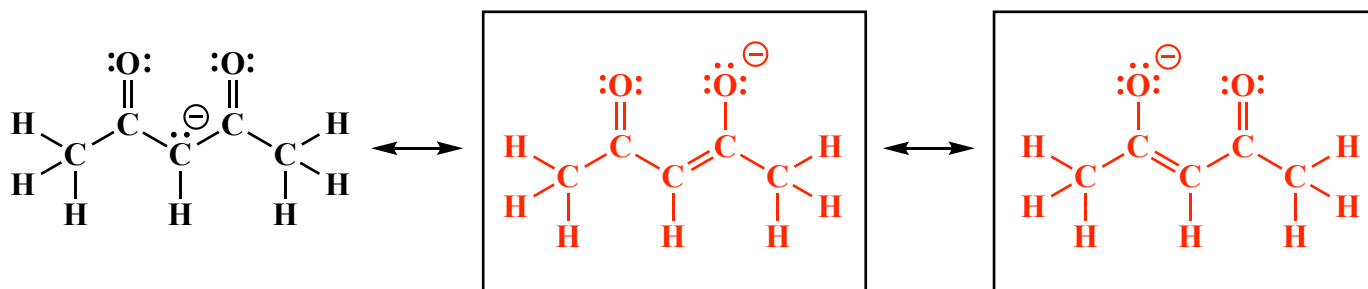
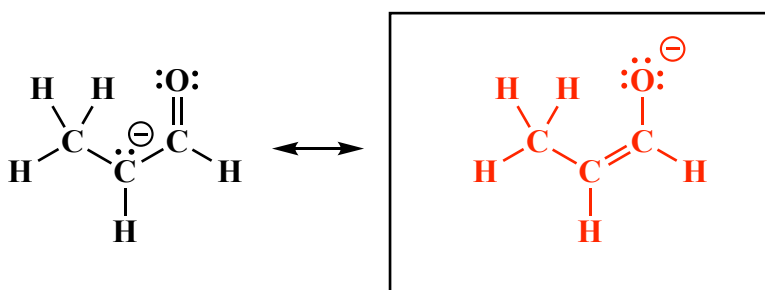
4. (6 pts each) Write an acceptable IUPAC name or draw a structural formula for the following molecules:

In the box, draw the structure corresponding to the following IUPAC name.

(E)-3-Ethyl-5-oxohex-2-enamide
or (E)-3-Ethyl-5-oxo-2-hexenamide

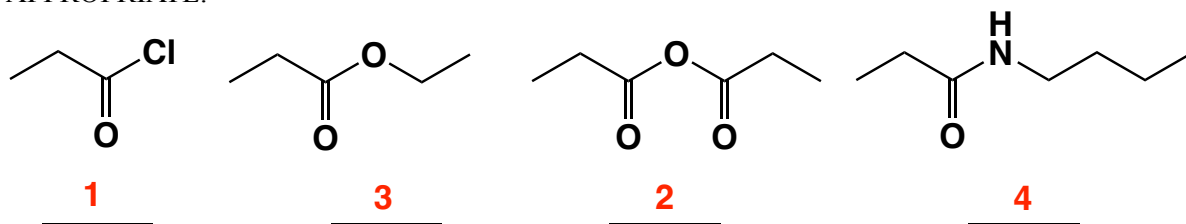


5. (9 pts each) For the two different enolates shown below, draw the other important contributing structures. Make sure to show all electrons and formal charges.

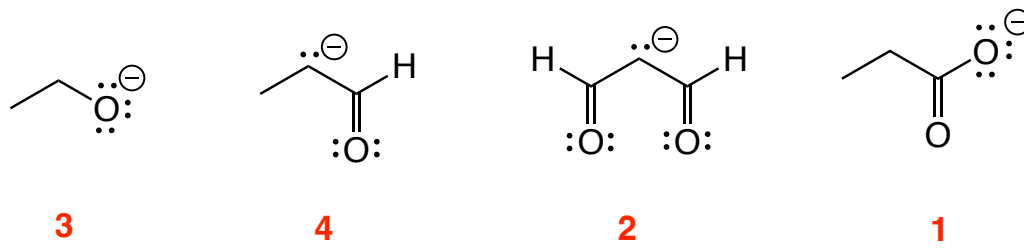


6. (14 pts) These are the ranking questions.

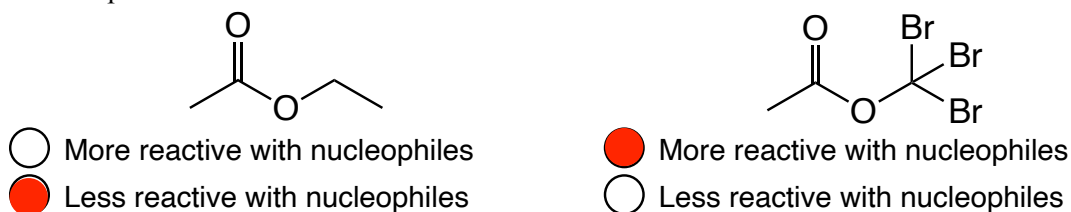
A) Rank the following with respect to reactivity with nucleophiles, WITH A "1" UNDER THE MOST REACTIVE AND "4" UNDER THE LEAST REACTIVE, AND THEN "2" AND "3" AS APPROPRIATE.



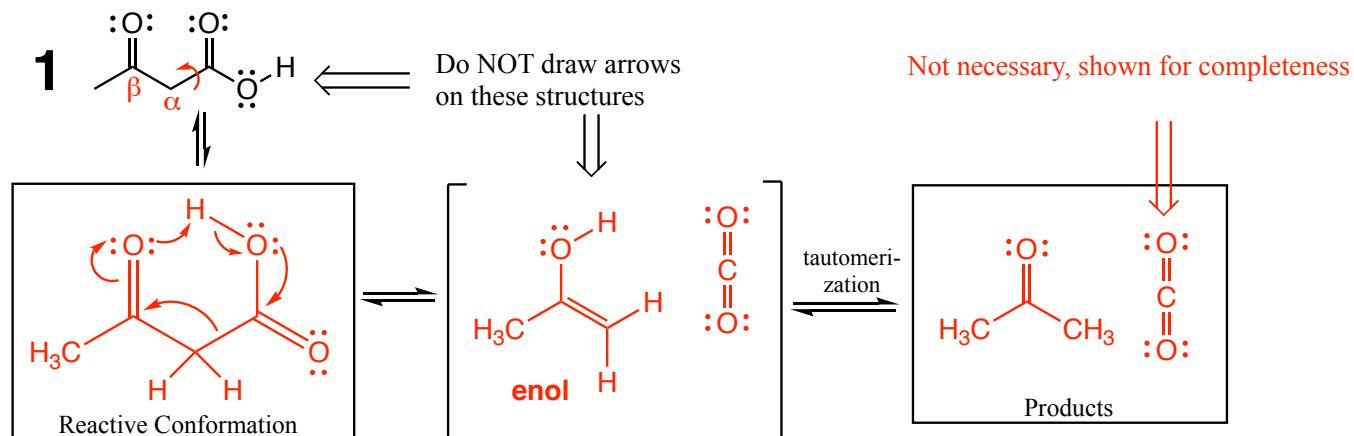
B) Rank the following with respect to anion stability, WITH A "1" UNDER THE MOST STABLE ANION AND "4" UNDER THE LEAST STABLE ANION, AND THEN "2" AND "3" AS APPROPRIATE.



C) For each pair of molecules, fill in the circles to indicate which in each pair is more or less reactive with nucleophiles.



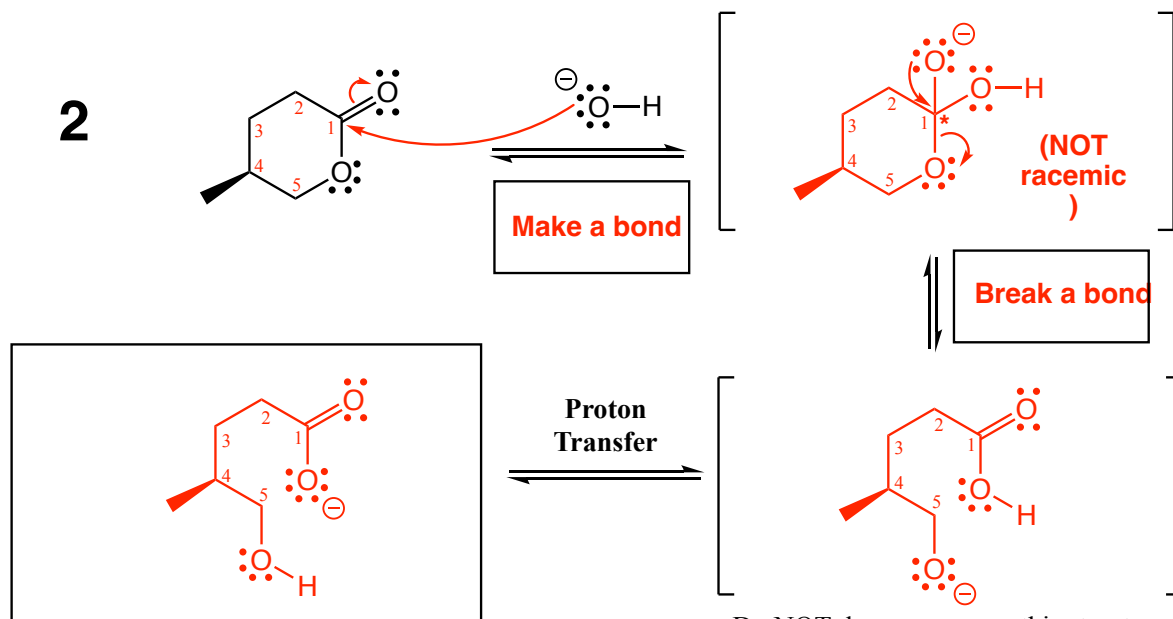
7. (12 pts) Complete the mechanism for the following decarboxylation reaction. **Be sure to show arrows to indicate movement of all electrons on the “Reactive Conformation”, write all lone pairs, all formal charges, and all the products for each step.** Remember, I said all the products for each step. **IF A NEW CHIRAL CENTER IS CREATED IN AN INTERMEDIATE OR PRODUCT, MARK IT WITH AN ASTERISK AND LABEL THE MOLECULE AS RACEMIC IF APPROPRIATE.**



Draw arrows on this structure

Note you will have to write a balanced equation for the above mechanism on page 7

8. (12 pts) Complete the mechanism for the following reaction of a lactone and hydroxide. **Be sure to show arrows to indicate movement of all electrons, write all lone pairs, all formal charges, and all the products for each step.** Remember, I said all the products for each step. **IF A NEW CHIRAL CENTER IS CREATED IN AN INTERMEDIATE OR PRODUCT, MARK IT WITH AN ASTERISK AND LABEL THE MOLECULE AS RACEMIC IF APPROPRIATE.** In the boxes provided, write which of the 4 mechanistic elements describes each step (make a bond, break a bond, etc.).



Do NOT draw arrows on this structure

Note you will have to write a balanced equation for the above mechanism on the page 7

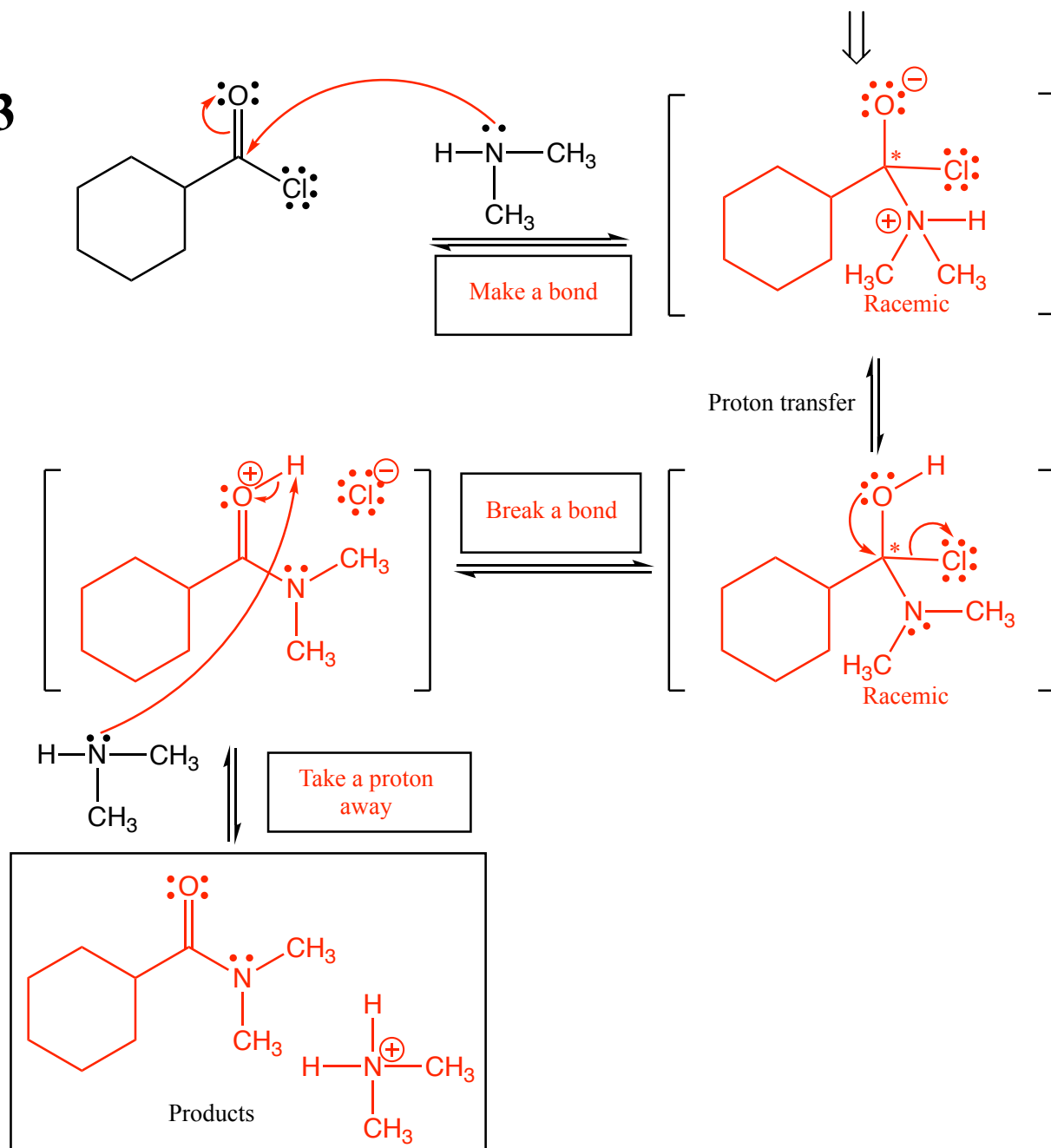
Signature _____

Pg 5 _____ (19)

9. (19 pts) Complete the mechanism for the following reaction of an acid chloride with an amine. **Be sure to show arrows to indicate movement of all electrons, write all lone pairs, all formal charges, and all the products for each step.** Remember, I said all the products for each step. **IF A NEW CHIRAL CENTER IS CREATED IN AN INTERMEDIATE OR PRODUCT, MARK IT WITH AN ASTERISK AND LABEL THE MOLECULE AS RACEMIC IF APPROPRIATE.** In the boxes provided, write which of the 4 mechanistic elements describes each step (make a bond, break a bond, etc.).

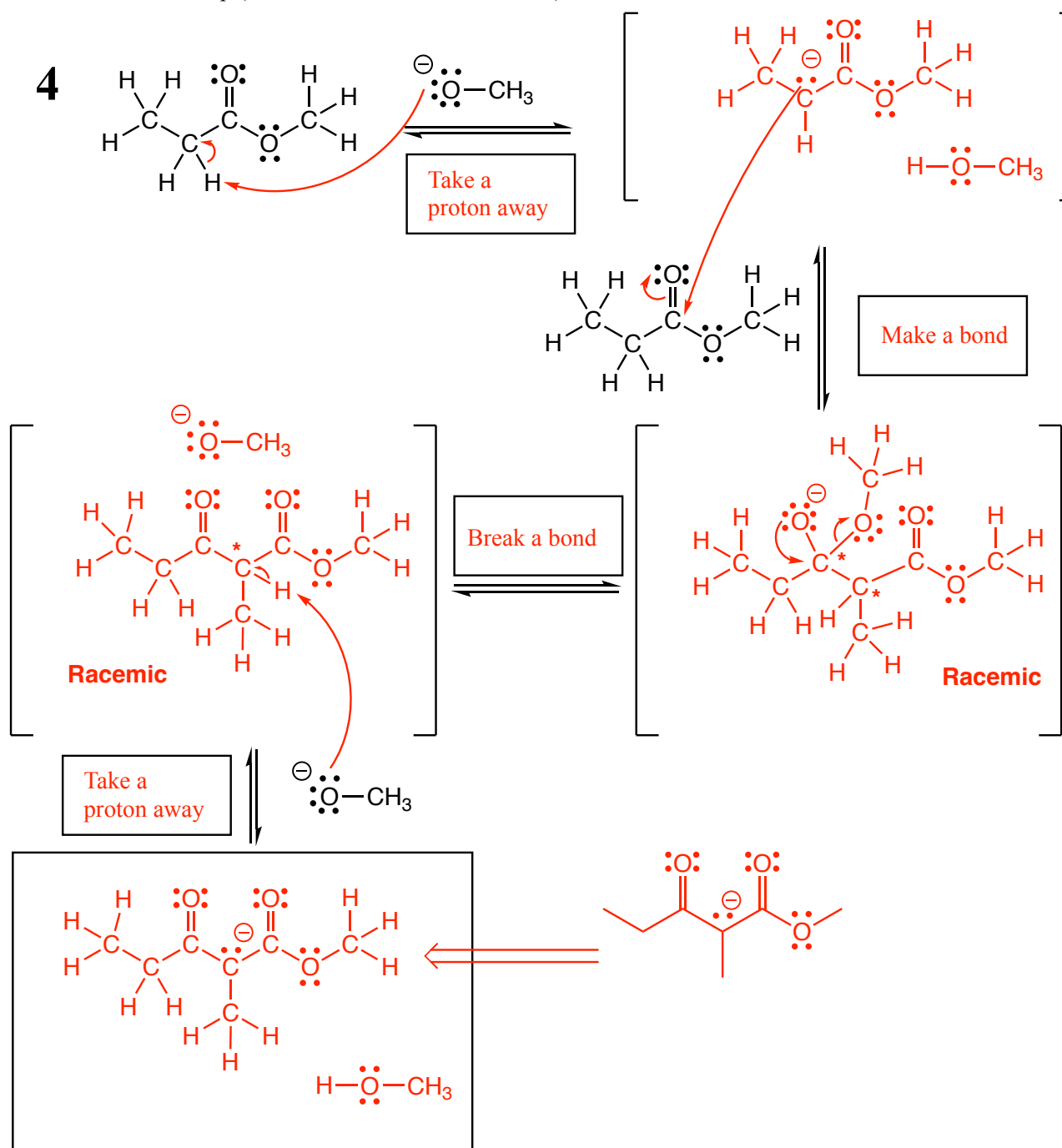
Don't write arrows on this structure

3



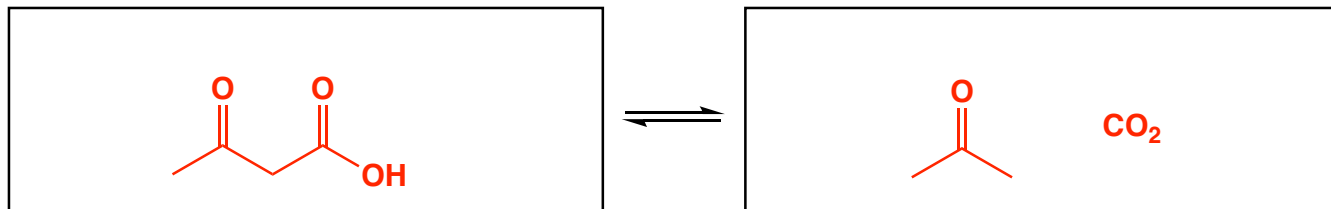
Note you will have to write a balanced equation for the above mechanism on PAGE 7

10. (23 pts) Complete the mechanism for the following Claisen condensation reaction. Be sure to show arrows to indicate movement of all electrons, write all lone pairs, all formal charges, and all the products for each step. Remember, I said all the products for each step. IF A NEW CHIRAL CENTER IS CREATED IN AN INTERMEDIATE OR PRODUCT, MARK IT WITH AN ASTERISK AND LABEL THE MOLECULE AS RACEMIC IF APPROPRIATE. In the boxes provided, write which of the 4 mechanistic elements describes each step (make a bond, break a bond, etc.).

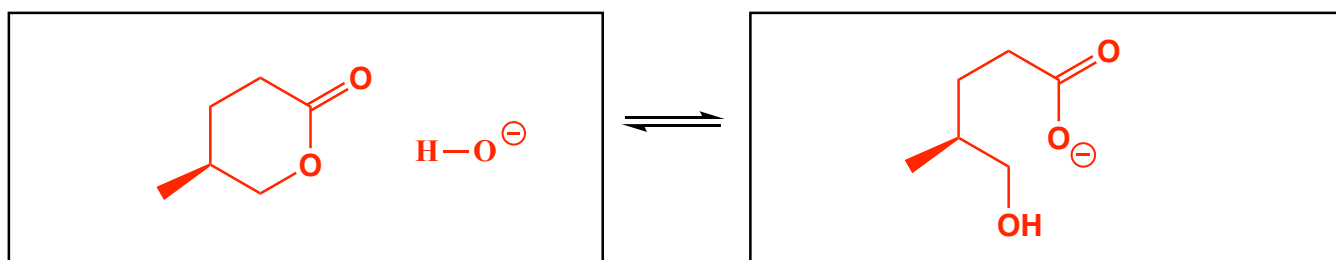


11. (17 pts) Write BALANCED equations for the four mechanisms, 1-4, that you drew on the last three pages. Only include molecules consumed or created during the reactions. In addition, you must use whole numbers when designating stoichiometries, not fractions or decimals. This is not asking to give equivalents, but rather balanced equations for each reaction.

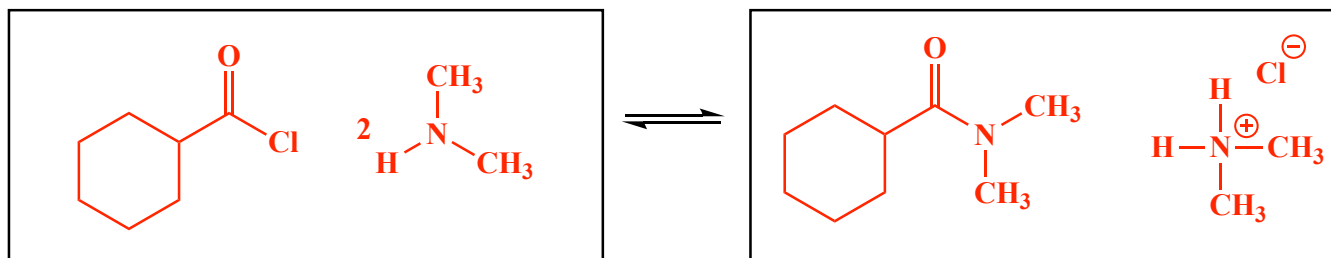
Write a balanced equation for the overall process described by mechanism 1 from page 4



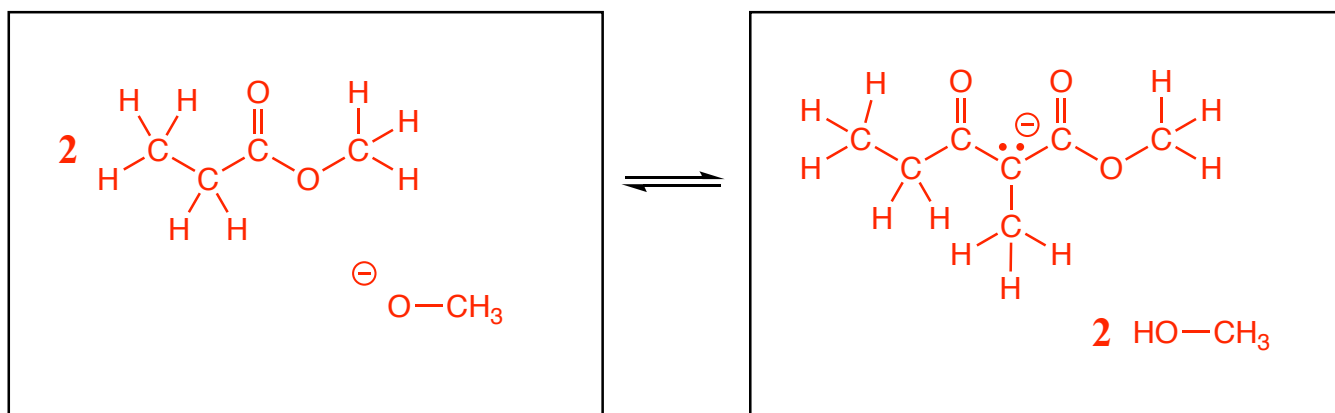
Write a balanced equation for the overall process described by mechanism 2 from page 4



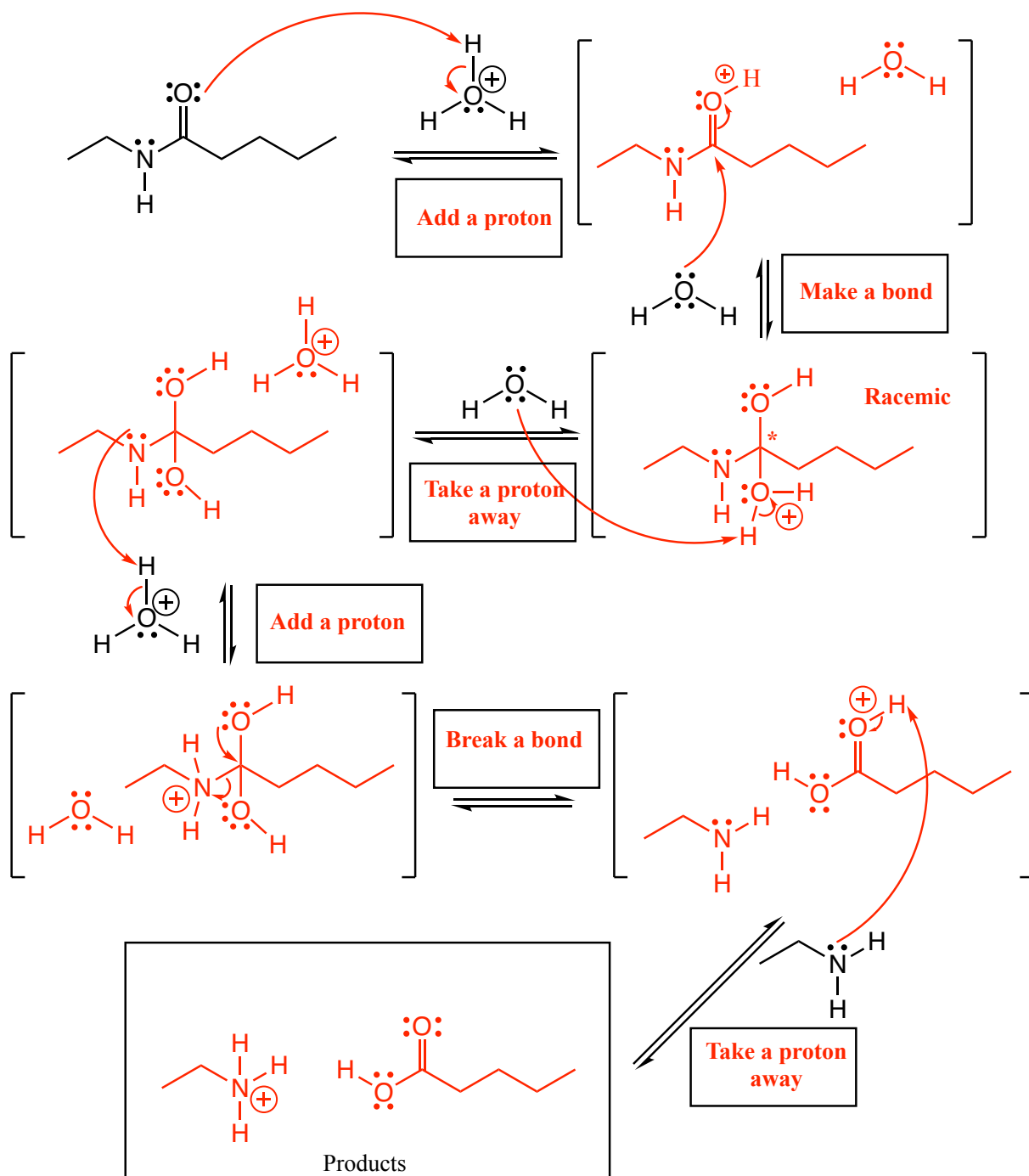
Write a balanced equation for the overall process described by mechanism 3 from page 5



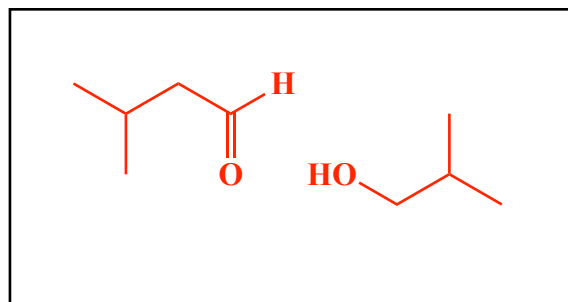
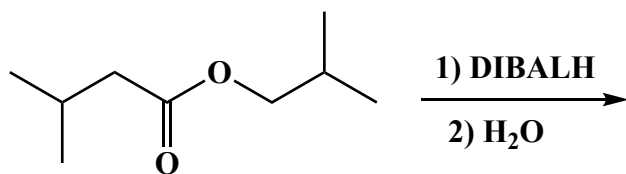
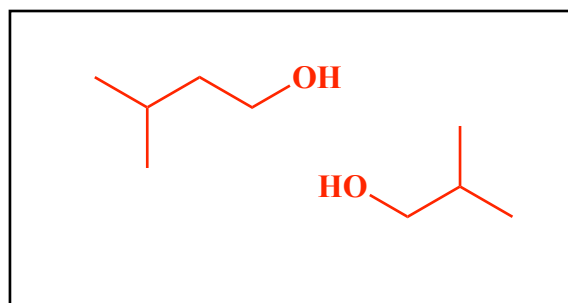
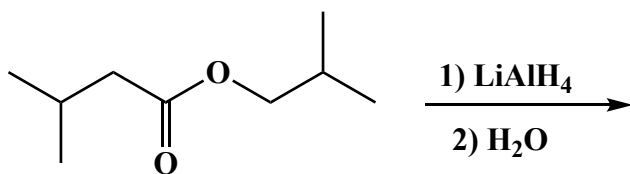
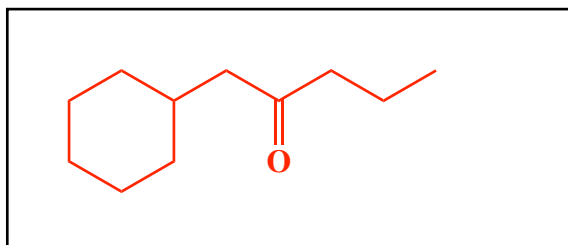
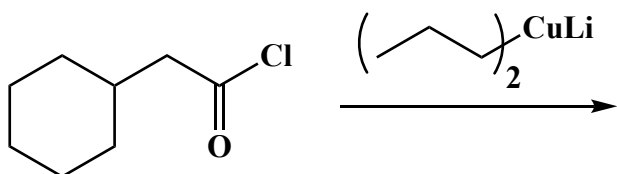
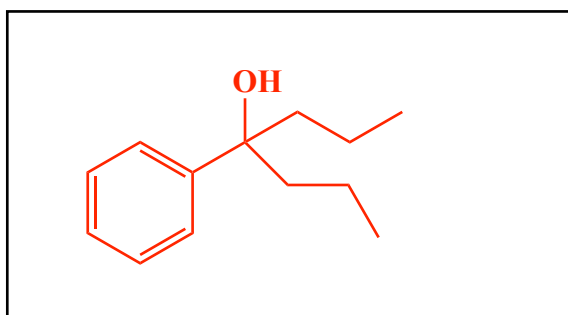
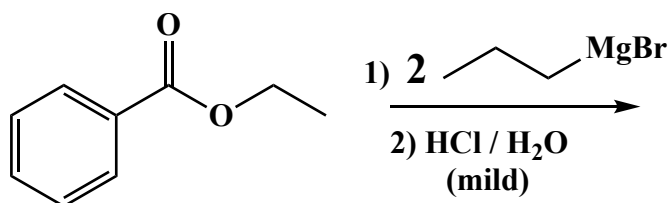
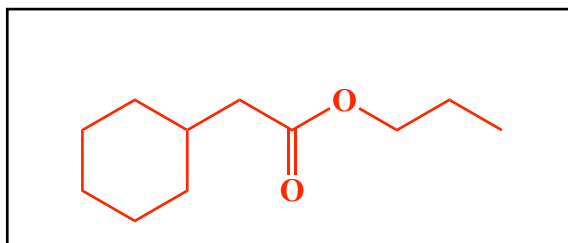
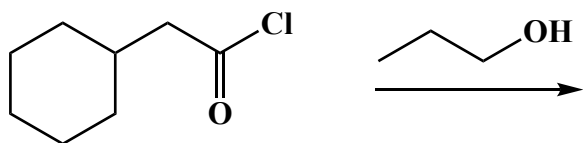
Write a balanced equation for the overall process described by mechanism 4 from page 6



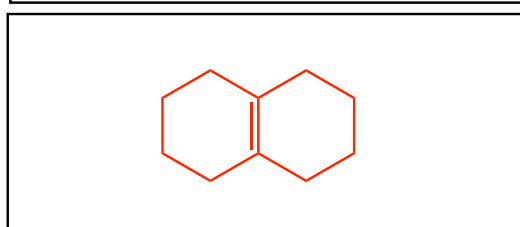
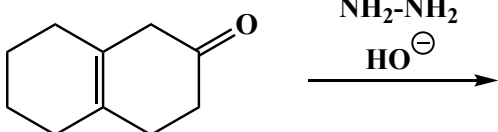
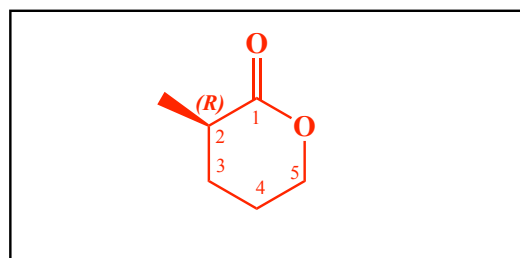
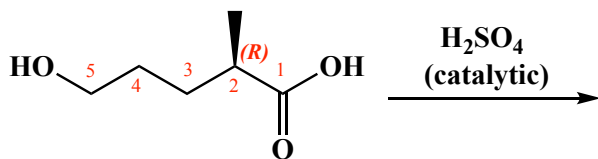
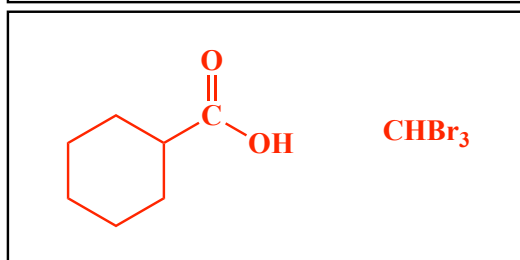
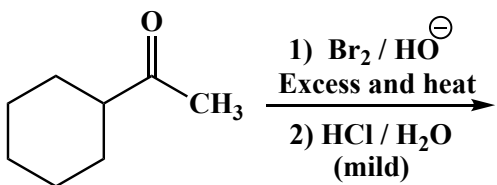
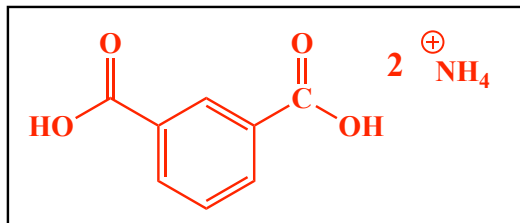
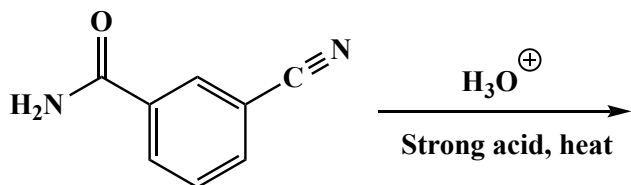
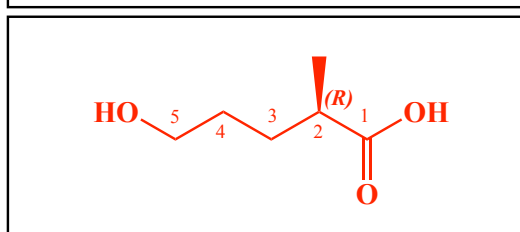
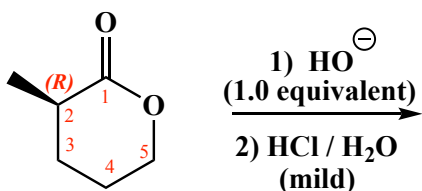
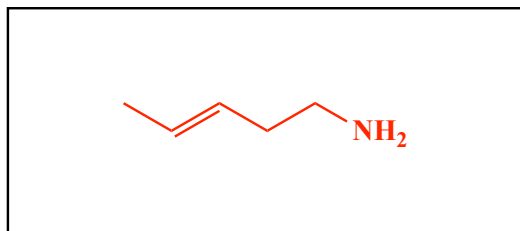
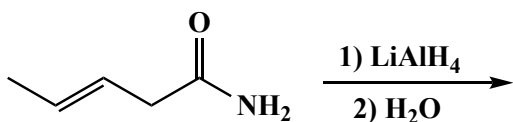
12. (35 pts) For this acid promoted amide hydrolysis reaction, use **arrows to indicate movement of all electrons**, write **all lone pairs**, **all formal charges**, and **all the products for each step**. **IF A NEW CHIRAL CENTER IS CREATED IN AN INTERMEDIATE, MARK IT WITH AN ASTERISK AND LABEL THE MOLECULE AS "RACEMIC" IF APPROPRIATE. FOR ALL CHIRAL PRODUCTS YOU MUST DRAW ALL ENANTIOMERS WITH WEDGES AND DASHES AND WRITE "RACEMIC" IF APPROPRIATE.** In the boxes provided by the arrows, write which of the 4 most common mechanistic elements describes each step (make a bond, break a bond, etc.).



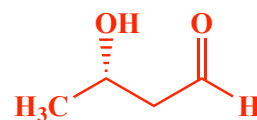
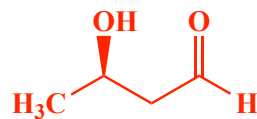
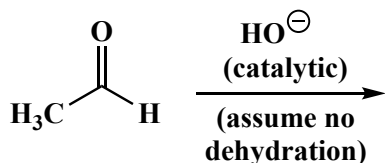
13. (3 or 5 pts.) Write all of the organic product(s) that will occur for each transformation. If a new chiral center is created and a racemic mixture is formed, you must draw both enantiomers and write "racemic" under the structure. Use wedges (\blacktriangleleft) and dashes (\cdots) to indicate stereochemistry. **For these, you need to write all of the products of the reactions except for the products containing metals.**



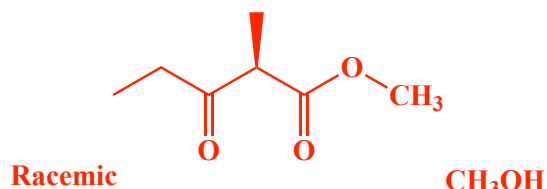
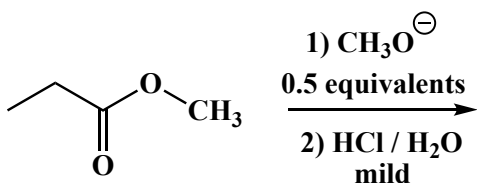
14. (3 or 5 pts.) Write all of the organic product(s) that will occur for each transformation. If a new chiral center is created and a racemic mixture is formed, you must draw both enantiomers and write "racemic" under the structure. Use wedges (\blacktriangleleft) and dashes (\cdots) to indicate stereochemistry. **For these, you need to write all of the products of the reactions except for the products containing metals.**



15. (5 or 9 pts.) Write all of the organic product(s) that will occur for each transformation. If a new chiral center is created and a racemic mixture is formed, you must draw both enantiomers and write "racemic" under the structure. Use wedges (\blacktriangle) and dashes (\cdots) to indicate stereochemistry. **For these, you need to write all of the products of the reactions except for the products containing metals.**

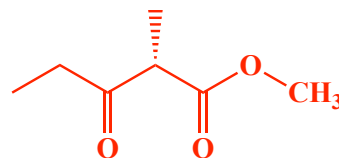


Racemic

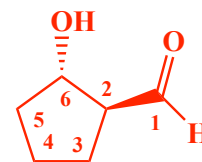
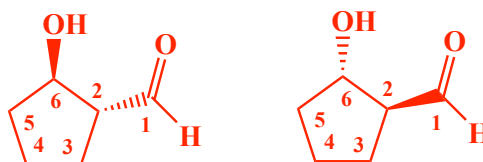
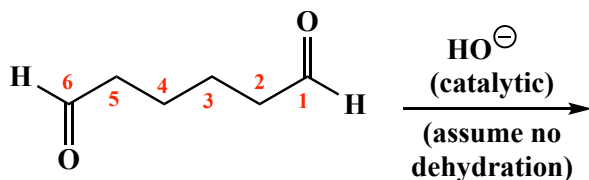


Racemic

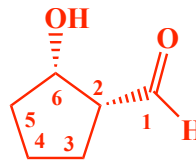
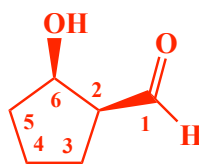
CH₃OH



This is a hard one!!



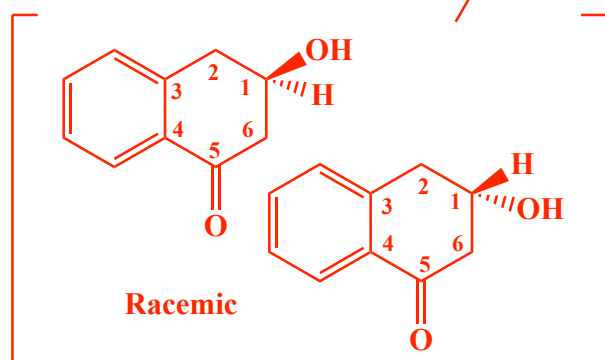
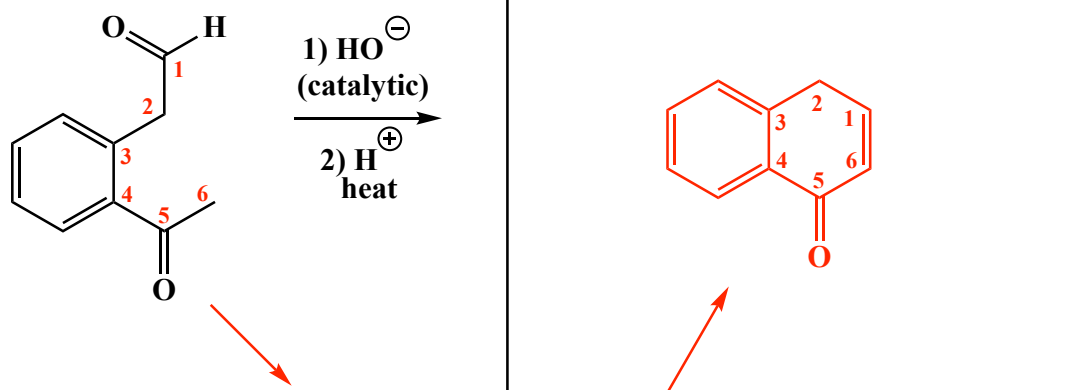
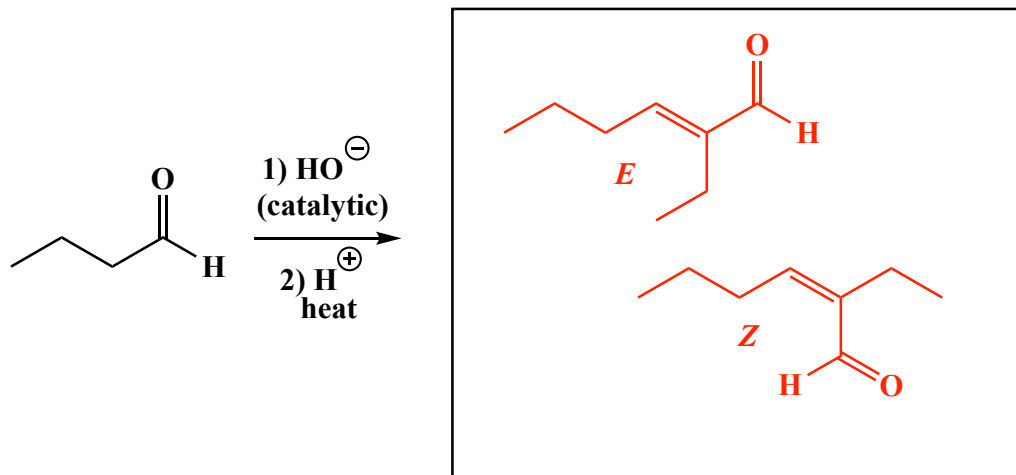
Racemic



Hint: there are four stereoisomers here

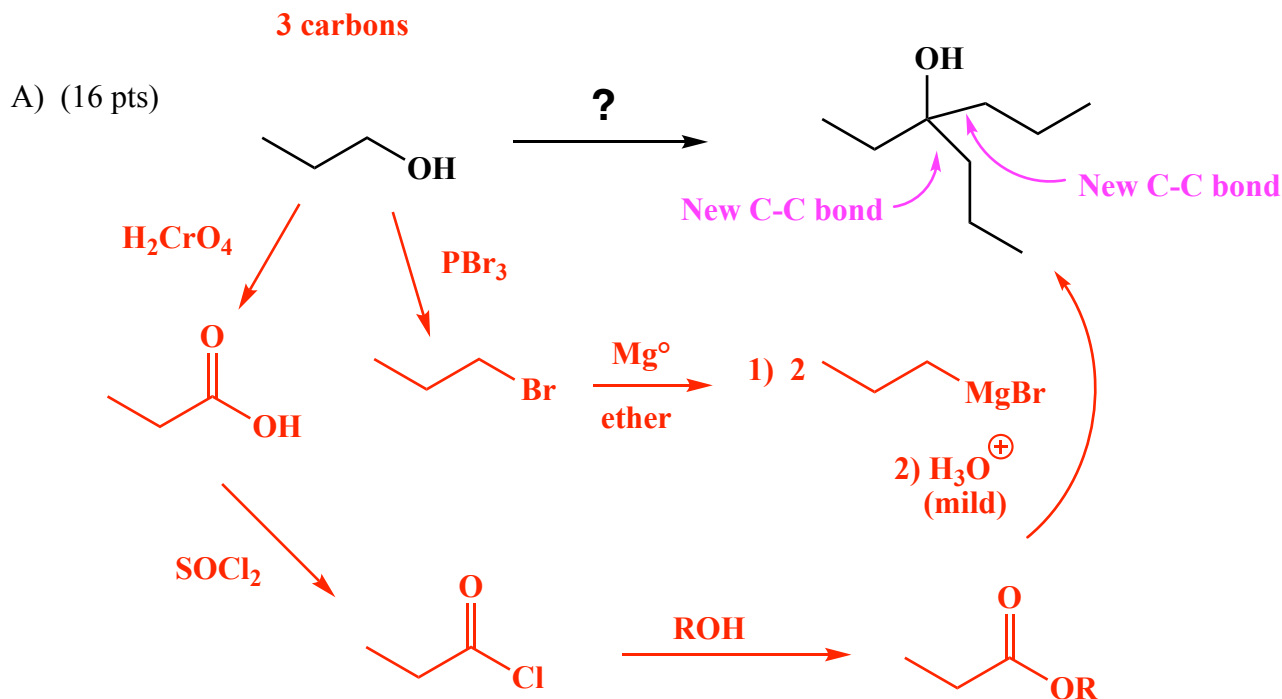
16. (4 or 6 pts.) Write the predominant product that will occur for each transformation. If a new chiral center is created and a racemic mixture is formed, you must draw both enantiomers and write "racemic" under the structure. Use wedges (\blacktriangleleft) and dashes (\cdots) to indicate stereochemistry. **For these, you need to write all of the products of the reactions except for the products containing metals.**

There is a lot to think about here. Please take your time. ASSUME THESE DEHYDRATES.



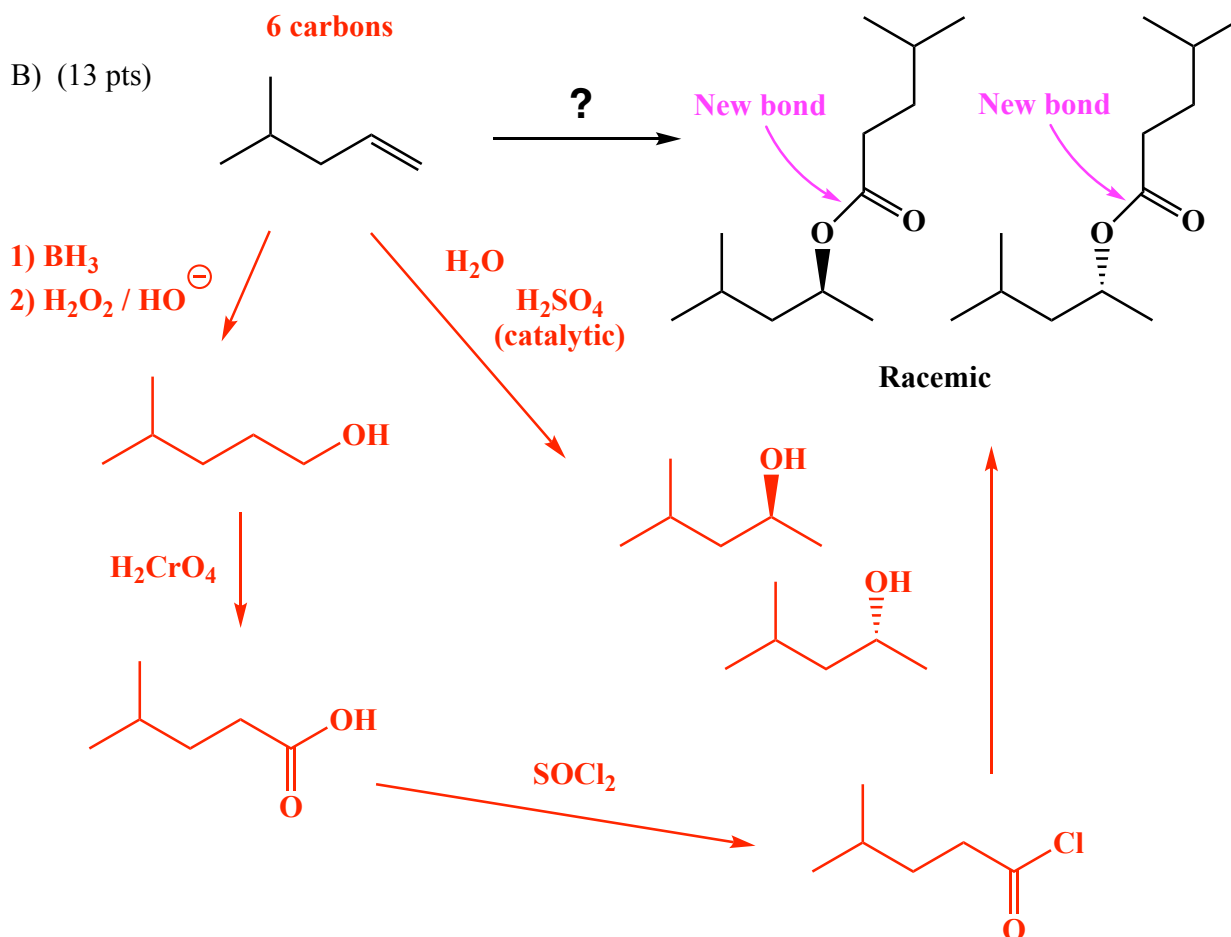
17. These are synthesis questions. You need to show how the starting material can be converted into the product(s) shown. You may use any reactions we have learned provided that the product(s) you draw for each step is/are the predominant one(s). Show all the reagents you need. Show each molecule synthesized along the way and be sure to pay attention to the regiochemistry and stereochemistry preferences for each reaction. You must draw all stereoisomers formed, and use wedges and dashes to indicate chirality at each chiral center. Write racemic when appropriate. **All the carbons of the product must come from carbons of the starting material.**

9 carbons



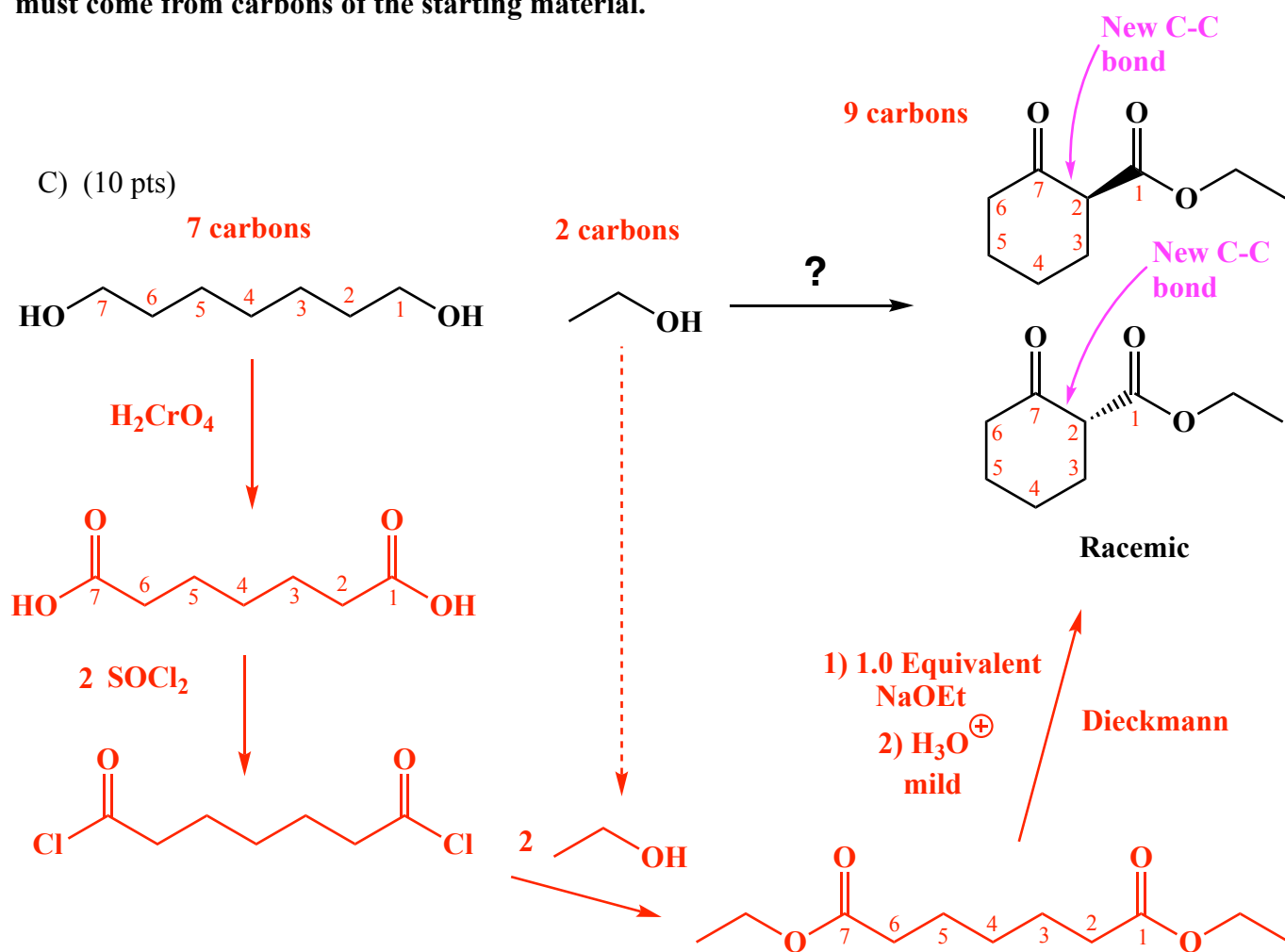
Recognize that the product is an ester with 9 carbon atoms, and the starting material has 3 carbon atoms. Therefore, three starting material molecules will be combined in the product. **Recognize** also that the product is a tertiary alcohol with two new carbon-carbon bonds on the same carbon as the OH group, and the two new bonds are made to two identical groups. That is the KRE of a Grignard reacting with an ester! Therefore, propose that the last step is the ester shown reacting with two equivalents of a three-carbon Grignard reagent. **Recognize** that the required Grignard reagent can be made from the starting alcohol through a reaction with PBr_3 followed by reaction with Mg° in ether. **Recognize** further that the required ester can be made from the starting alcohol through the sequence of oxidation with H_2CrO_4 to give the carboxylic acid, followed by reaction with SOCl_2 to give the acid chloride that reacts with any alcohol to give the ester. Note we do not care which ROH is used to make the ester because those carbons do not end up in the product.

17. These are synthesis questions. You need to show how the starting material can be converted into the product(s) shown. You may use any reactions we have learned provided that the product(s) you draw for each step is/are the predominant one(s). Show all the reagents you need. Show each molecule synthesized along the way and be sure to pay attention to the regiochemistry and stereochemistry preferences for each reaction. You must draw all stereoisomers formed, and use wedges and dashes to indicate chirality at each chiral center. Write racemic when appropriate. **All the carbons of the product must come from carbons of the starting material.** **12 carbons**



Recognize that the product is an ester with 12 carbon atoms, and the starting material has 6 carbon atoms. Therefore, two starting material molecules will be combined in the product. **Recognize** also that the C-O single bond next to the carbonyl of an ester is the one that can be made. Therefore, the last reaction is between an acid chloride and an alcohol as shown (also could be Fischer esterification using a carboxylic acid and an alcohol with catalytic H_2SO_4). The alcohol can be made from the starting alkene using hydroboration/oxidation ($\text{BH}_3 / \text{H}_2\text{O}_2$ and HO^\ominus), and the acid can be made from oxidizing that alcohol with H_2CrO_4 . The acid chloride is made from the acid using SOCl_2 . **Recognize** that the secondary alcohol needed to make the ester is made from the starting alkene in one step using the acid-catalyzed hydration reaction that gives Markovnikov regiochemistry as shown. Note, in this case we will not want you to use the Fischer esterification reaction because of competition with the dehydration of the secondary alcohol if H_2SO_4 were added.

17. These are synthesis questions. You need to show how the starting material can be converted into the product(s) shown. You may use any reactions we have learned provided that the product(s) you draw for each step is/are the predominant one(s). Show all the reagents you need. Show each molecule synthesized along the way and be sure to pay attention to the regiochemistry and stereochemistry preferences for each reaction. You must draw all stereoisomers formed, and use wedges and dashes to indicate chirality at each chiral center. Write racemic when appropriate. **All the carbons of the product must come from carbons of the starting material.**



Recognize that the product is a cyclic β -keto ester with 9 carbon atoms, and the starting materials have 7 carbon atoms and 2 carbon atoms, respectively. **Recognize** further that a cyclic β -keto ester is the KRE of a Dieckmann condensation with a new C-C bond as shown. **Recognize** that the required diester needed for the Dieckmann condensation is a diethyl ester, so assume the ethanol starting material is used to create the ester from the 7-carbon diacid chloride as shown. **Recognize** that the required 7-carbon diacid chloride can be made from starting 7-carbon diol from the sequence of reacting with H_2CrO_4 to give the diacid followed by reaction with SOCl_2 to give the diacid chloride. Note that it would have been fine to react the 7-carbon diacid with ethanol in the presence of catalytic H_2SO_4 (Fischer esterification) instead of making the diacid chloride.

I hope you all have a wonderful spring break. Please make a promise to yourself to take some time to do things you really enjoy. **YOU DESERVE IT**, after all, you are in OChem II! And, of course, all of next week make sure to **EXERCISE EVERY CHANCE YOU GET**. Our 3.1 mile challenge is coming up the first week of April!